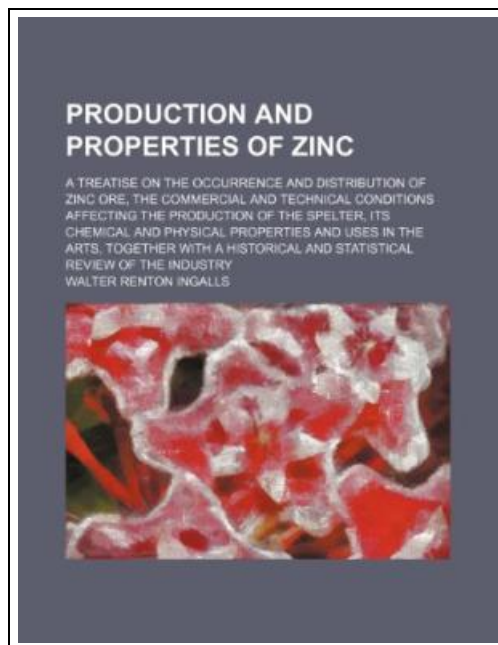


## Production and Properties of Zinc; A Treatise on the Occurrence and Distribution of Zinc Ore, the Commercial and Technical Conditions Affecting the Production of the Spelter, Its Chemical and Physical Properties and Uses



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1902 Excerpt: . disengagement of carbonic acid and the precipitate formed is a basic hydrocarbonate. This hydrocarbonate is soluble in an aqueous solution of carbonic acid; upon exposure to the air a granular powder deposits from the solution, which has been taken for neutral carbonate of zinc by some chemists, but by others has been held to be a basic carbonate. According to Thorpe<sup>2</sup> neutral zinc carbonate,  $\text{ZnCO}_3$ , can be produced artificially by adding acid carbonate of sodium to a solution of zinc sulphate, while Roscoe and Sehorellemmer state that it is produced by precipitation from  $\text{ZnSO}_4$  with an excess of hydrogen potassium carbonate. In a paper on Sulphide Ore Treatment, Metallurgy, London, June 19, 1901. read before the Institution of Mining and 2 Dictionary of Applied Chemistry, p. 1057. bicarbonate ( $\text{KHCO}_3$ ), the neutral carbonate of sodium ( $\text{Na}_2\text{CO}_3$ ) throwing down a basic carbonate of zinc, which is the more basic the higher the temperature of precipitation and the more dilute the solution. Zinc carbonate,  $\text{ZnCO}_3$ , is produced when the solution of a zinc salt is precipitated by an alkaline carbonate under pressure. As obtained by heating a solution of  $\text{ZnCl}_2$  to 150-160 C. and precipitating with calcium carbonate or sodium bicarbonate it is a white microcrystalline powder. The neutral hydrated carbonate ( $2\text{ZnCO}_3 \cdot \text{H}_2\text{O}$ ) is an amorphous powder obtained by digesting a basic carbonate of zinc with ammonium bicarbonate. The anhydrous carbonate of zinc is soluble in carbonic acid water. A solution of  $\text{CO}_2$  at five atmospheres, taking...



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